**Chemistry 11 – Unit 5 —Solution Chemistry**

**Unit Outline 2015**

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| **Topic** | **Reference** |
| Definitions and Types of SolutionsIndividual Ion ConcentrationsPrecipitation ReactionsTitrationsTest | Definitions of solute, solvent, solution, saturated, unsaturated and solubility.Definition of molarity and ability to calculate from grams to moles to molarity, etc...Polar covalent and non-polar covalent solvents and ionic, polar covalent and non-polar covalent solutes.Be able to describe what happens to ions when they dissolve in water (see class notes and p. 373)**Reference: Problem Set 5.1, Text p. 364-368**Dissociation and concentrations of individual ions in solution.Calculations involving dilution, individual ion concentrations and mixtures of solutions. (you will NOT be tested on electrical conductivity)**Reference: Text p. 383-388, Problem Set 5.2, 5.3 and 5.4**Determining solubility - **Solubility chart in data booklet.**Precipitation Reactions and Writing Formula, Complete Ionic and Net Ionic Equations for Precipitation Reactions. Calculation of ions in solution after a precipitation reaction**Reference: Problem Set 5.5, 5.6 and Lab 5.2****Reference: Text p. 393, 395-396, Titrations worksheet on weebly** Thursday, February 12, 2015 |